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Introduction to transport equations - Part

1 Transport equation

Introduction to transport equations - Part 2 PDE-3 | Transport equation: derivation PDE-4 | Transport equation: general solution [Deriving a conservation equation](#) The transport equation *Introduction to the Boltzmann transport equation (BTE) PGE 381M Lecture 2.1*

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Thus, the proper simulation of flows in rarefied gases requires a more detailed description. This book discusses classical and modern methods to derive macroscopic transport equations for rarefied gases from the Boltzmann equation, for small and moderate Knudsen numbers, i.e. at and above the Navier-Stokes-Fourier level.

Macroscopic Transport Equations for Rarefied Gas Flows ...

8.5.6 Comparison with Jin-Slemrod equations 143 9 Macroscopic transport

equations for rarefied gas flows 145 9.1 Relations between the equations 145 9.2 3-D non-linear equations 146 9.2.1 Conservation laws 146 9.2.2 Chapman-Enskog expansion 147 9.2.3 Moment equations for Maxwell molecules 150 9.2.4 Moment equations for general molecule types 152

Macroscopic Transport Equations for Rarefied Gas Flows

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MACROSCOPIC TRANSPORT MODELS FOR RAREFIED GAS FLOWS7 of 26 The idea of the CE expansion method is to add corrections to the local equilibrium distribution by adding terms of higher orders in the Knudsen number, $f = f(0) + Kn f(1) + Kn^2 f(2) + \dots$, (3.2) subject to the condition that the hydrodynamic variables $\{\rho, v_i, \theta\}$ are the same at any level of expansion, so that $\rho = \rho^{(0)}, v_i = v_i^{(0)}, \theta = \theta^{(0)}$

Macroscopic transport models for rarefied gas flows: a ...

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Couette flows from low speed to high speed, with particular concentration on the detailed structure of the nonisothermal KL and the shear-stress Knudsen number dependence of the effective transport coefficients in the whole system.

Nonlinear transport of rarefied Couette flows from low ...

Many macroscopic equations are proposed to describe the rarefied gas dynamics beyond the Navier-Stokes level, either from the mesoscopic Boltzmann equation or some physical arguments, including (i) Burnett, Woods, super-Burnett, augmented Burnett equations derived from the Chapman-Enskog expansion of the Boltzmann equation, (ii) Grad 13, regularized 13/26 moment equations, rational extended thermodynamics equations, and generalized hydrodynamic equations, where the velocity distribution ...

On the accuracy of macroscopic equations for linearized ...

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Macroscopic and kinetic modelling of rarefied polyatomic ...

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Thermophoresis of a spherical particle:

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Influence of angular momentum on transport coefficients in ...

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Macroscopic Transport Equations For Rarefied Gas Flows ...

The kinetic theory of gases is a historically significant, but simple, model of the thermodynamic behavior of gases, with which many principal concepts of thermodynamics were established. The model describes a gas as a large number of identical submicroscopic particles (atoms or molecules), all of which are in constant, rapid, random motion. Their size is assumed to be much smaller than the ...

8.5.6 Comparison with Jin-Slemrod equations 143
 9 Macroscopic transport equations for rarefied gas flows 145
 9.1 Relations between the equations 145
 9.2 3-D non-linear equations 146
 9.2.1 Conservation laws 146
 9.2.2 Chapman-Enskog expansion 147
 9.2.3 Moment equations for Maxwell molecules 150
 9.2.4 Moment equations for general molecule types 152

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